

Remediation of Dicofol Type Ddts-Contaminated Sediments by Ferrous Activated Sodium Persulfate Oxidation

Chun-You Zhu¹, Peng Bao², Yu-Xin Ba¹, Jing Hua¹, Xiao-Ning Liu¹, Guo-Hua Hou², Chun-Zao Liu² & Zheng-Yi Hu¹

¹ College of Resources and Environment, University of Chinese Academy of Sciences, Beijing, P. R. China

² State Key Lab of Urban and Regional Ecology, Research Center for Eco-Environmental Sciences, Chinese Academy of Sciences, Beijing, P. R. China

Correspondence: Zheng-Yi Hu, College of Resources and Environment, University of Chinese Academy of Sciences, 19A, Yuquan Road, Beijing 100049, China. Tel: 86-10-8825-6542. E-mail: zhyhu@ucas.ac.cn

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Abstract

In recent years, contamination by dicofol-type DDTs has attracted immense concern as a new source of DDT pollution. In this study, sediment samples from a dicofol manufacturing factory in Tianjing, China exhibited serious DDT contamination [*p,p'*-DDE (115.27 mg kg⁻¹) and *p,p'*-DDT (11.84 mg kg⁻¹)]. Results of the batch experiments showed that total DDT degradation rates increase as S₂O₈²⁻/Fe²⁺ molar ratios increase. The S₂O₈²⁻/Fe²⁺ molar ratios used in this study were as follows: 60/10 < 10/30 < 20/30 < 60/50 < 60/20 < 40/30 < 60/40 < 60/30 < 80/30. Their corresponding degradation rates were 31, 43, 52, 69, 70, 71, 72, 89, and 91 μg g⁻¹, respectively. The optimal S₂O₈²⁻/Fe²⁺ molar ratio was 60/30, which resulted in 64% and 96% degradation of *p,p'*-DDE and *p,p'*-DDT, respectively. However, when an excessive amount of ferrous ion was used (<S₂O₈²⁻/Fe²⁺ molar ratio of 60/30), then competition for SO₄⁻ between ferrous ion and DDTs resulted in decreased DDT degradation efficiency and increased persulfate decomposition (represented by the generated amount of sulfate). Our results implied that a slow and steady production of sulfate free radicals is favorable for DDT degradation, and that Fe²⁺ availability plays an important role in controlling persulfate reactions activated by ferrous ion. Fe²⁺-activated persulfate oxidation may be significant in developing environment friendly and fast-remediation options for DDT-contaminated sediments and soil. Therefore, this study contributes to current knowledge on remediating DDT contamination.

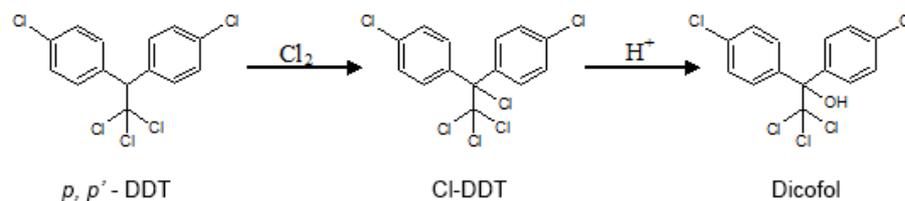
Keywords: sediments, dicofol-type DDT contamination, persulfate, ferrous ion, oxidative degradation

1. Introduction

DDT [1,1,1-trichloro-2,2-bis(*p*-chlorophenyl)-ethane] is one of the persistent organic pollutants (POPs) identified by the Stockholm Convention on POPs which has been extensively used for controlling agricultural pests and disease-carrying insects such as malaria vectors (Zitko, 2003; Kamanavalli & Ninnekar, 2004). DDT is more stable than other organochlorine pollutants because of its chlorinated aliphatic and aromatic structures. Exposure to DDTs (DDT and its homologues) may damage the human nervous and reproductive systems (Guo et al., 2009). Although the manufacture and application of DDTs have been restricted since the 1970s because of their negative effects, traces of DDTs can still be detected in air, water, soil, sediments, and organisms (Bettinetti et al., 2008; Yao et al., 2006). Moreover, the 2001 Stockholm Convention on POPs still allows the use of DDTs in several countries, such as South Africa, to control the transmission of malaria. DDE [1-Chloro-2-[2,2-dichloro-1-(4-chlorophenyl)ethenyl]-benzene] is a common metabolite of DDT (Ssebugere et al., 2010; Yang et al., 2010). Environmental DDE originates from the metabolites of DDT resulting from aerobic biotic, abiotic, and photochemical degradation, as well as from technical-grade DDT contaminants (Thomas et al., 2008). DDE has been reported to be more persistent than DDT and can be detected in soil decades after the application of DDT (Thomas et al., 2008). According to the United States Geological Survey (USGS), *p,p'*-DDE content in America was 60% in urban areas and 48% in rural areas in 1999 (Thomas et al., 2008). As a potent androgen antagonist (Kelce et al., 1995), DDE has also been found to be the most abundant DDT component in sediments (Eganhouse & Pontolillo, 2008), fish, and humans (Kamanavalli & Ninnekar, 2004).

Dicofol [2,2,2-trichloro-1,1-bis(4-chlorophenyl)ethanol] is a non-systemic acaricide extensively used in

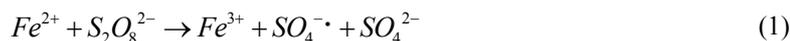
controlling mites. Dicofol is usually synthesized from technical *p,p'*-DDT. During synthesis reaction (Scheme 1), *p,p'*-DDT is first chlorinated into Cl-DDT, and then hydrolyzed into dicofol (Qiu et al., 2005).



Scheme 1. Synthesis reaction of dicofol by *p,p'*-DDT

To date, China still produces 5000 tons to 6000 tons of DDT per year as raw and processed materials for dicofol production (Huang et al., 2007). Moreover, approximately 8770 tons of DDTs were released into the environment in China by dicofol-type DDT contamination from 1988 to 2002 (Qiu et al., 2005; Turgut et al., 2009).

Although biodegradation and anaerobic reductive dechlorination for remediating DDTs have been well studied (Li et al., 2010; You et al., 1996), investigations on aerobic oxidative degradation for remediating DDTs is limited. Chemical oxidation via persulfate oxidation activated by ferrous ion has been evaluated as an option for treating chlorinated organic contaminants, such as trichloroethylene (TCE; Liang et al., 2004a, 2004b; Liang et al., 2008); tetrachloroethylene, dichloroethylene, and dichloroethane (Abranovic et al., 2006); polychlorinated biphenyls and polycyclic aromatic hydrocarbons (Block et al., 2004); and lindane (γ -HCH; Cao et al., 2008). Sulfate free radicals ($\text{SO}_4^{\cdot-}$) can be formed rapidly through persulfate-ferrous ion reaction at ambient temperature (20 °C) (Liang et al., 2004a). These free radicals can potentially degrade organic contaminants within soil mass by in situ chemical oxidation. The stoichiometric reaction between persulfate and ferrous ion is shown in the following equations (Kolthoff et al., 1951):



The ratio of reaction between $\text{S}_2\text{O}_8^{2-}$ and Fe^{2+} is dependent on the concentration of each reactant. When the reaction is near to stall, increasing the concentration of Fe^{2+} will accelerate the reaction shown in Equation (1). However, the target chlorinated organic contaminant and the excess Fe^{2+} will compete for $\text{SO}_4^{\cdot-}$, as shown in Equation (2). Gradual addition of small quantities of Fe^{2+} is necessary to optimize $\text{S}_2\text{O}_8^{2-}$ oxidative degradation of the target chlorinated organic contaminant and to control the reaction. To our knowledge, no Fe^{2+} -activated persulfate oxidation technique for remediating DDT contamination in sediments has yet been reported.

The primary purpose of this study is to investigate the contamination caused by dicofol-type DDTs in a dicofol manufacturing factory in Tianjin, China and to evaluate the effectiveness of Fe^{2+} -activated $\text{S}_2\text{O}_8^{2-}$ oxidative degradation for DDTs (*p,p'*-DDE and *p,p'*-DDT). In addition, the effects of various initial Fe^{2+} and $\text{S}_2\text{O}_8^{2-}$ concentrations on DDT degradation in sediments at ambient temperature (20 °C) are also investigated.

2. Materials and Methods

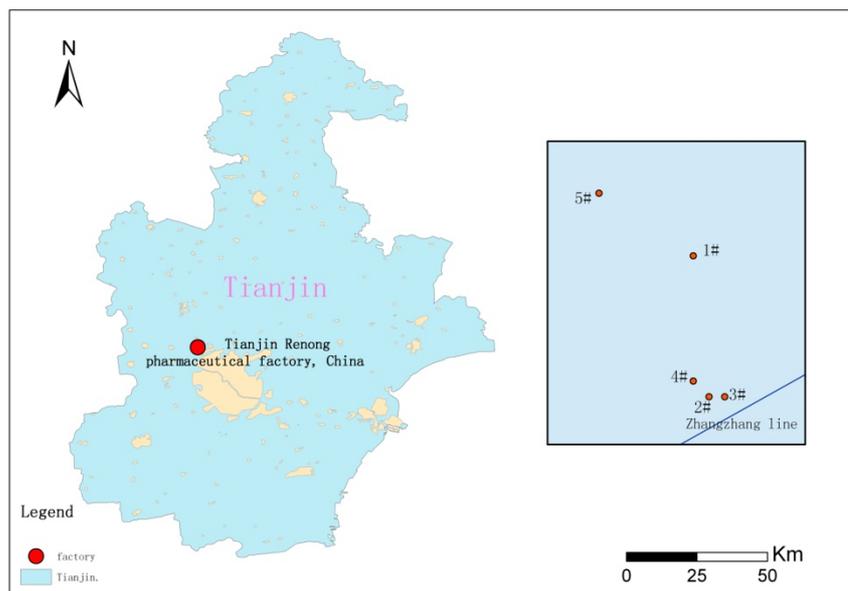
2.1 Chemicals

Standard samples of *p,p'*-DDT [*p,p'*-dichlorodiphenyltrichloroethane] (>99.5% purity) and *p,p'*-DDE [1-Chloro-2-[2,2-dichloro-1-(4-chlorophenyl)ethenyl]-benzene] (>99.5% purity) were obtained from Dr. Ehrenstorfer GmbH (Augsburg, Germany). Sodium persulfate ($\text{Na}_2\text{S}_2\text{O}_8$; >99% purity), ferrous sulfate ($\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$; >99% purity), anhydrous sodium sulfate (Na_2SO_4 ; >99% purity), and other chemical reagents were of analytical grade, as required. Super pure GC hexane was obtained from J&K Chemical Ltd. (China). Sodium persulfate and ferrous ion solutions were prepared with 18 M Ω deionized water (Milli-Q™ 18 M Ω system, Millipore Corporation, MA, USA) before use. All pieces of glassware were washed twice with hexane prior to use. Serum bottles (20 mL) were used as batch reactors.

2.2 Sediment Collection and Preparation

Sediment samples were collected from Tianjin Renong Pharmaceutical Factory in China (Scheme 2). Sampling sites were highly contaminated by DDTs because of dicofol manufacturing before 2002. The three sampling sites were the factory floor, sewage, and drainage ditch. Sediments were collected from 20 cm to 40 cm sections

mainly composed of fine sand with silt. The vessels were filled with sediments and completely sealed. The DDTs in the samples were detected via gas chromatography with mass selective detection (Table 1). Sample no. 3 was selected as the object for the batch study because it is highly contaminated with *p,p'*-DDE (115.27 mg kg⁻¹) and *p,p'*-DDT (11.84 mg kg⁻¹). The soluble sulfate of sample no. 3 was 56.28 mg kg⁻¹ and its pH was 7.7. In addition, small amounts of *p,p'*-DDD and *o,p'*-DDE were also detected in all samples. However, no attempt was made to quantify *p,p'*-DDD and *o,p'*-DDE because their concentrations were very low.



Scheme 2. The study area (Tianjin, China) and the sampling site (Tianjin Renong Pharmaceutical Factory)

2.3 Batch Studies of DDT Oxidation Using Persulfate Activated by Ferrous Ion

Two grams of air-dried sediment sample was added into the serum bottle with 4 mL deionized water and mixed completely. Two sets of experiments were performed to determine the following: (1) effect of $S_2O_8^{2-}$ concentration on DDT degradation and (2) effect of Fe^{2+} concentration on DDT degradation. Six treatments were initiated for set (1): (i) control, (ii) $Na_2S_2O_8$ (10 mmol L⁻¹) + $FeSO_4 \cdot 7H_2O$ (30 mmol L⁻¹), (iii) $Na_2S_2O_8$ (20 mmol L⁻¹) + $FeSO_4 \cdot 7H_2O$ (30 mmol L⁻¹), (iv) $Na_2S_2O_8$ (40 mmol L⁻¹) + $FeSO_4 \cdot 7H_2O$ (30 mmol L⁻¹), (v) $Na_2S_2O_8$ (60 mmol L⁻¹) + $FeSO_4 \cdot 7H_2O$ (30 mmol L⁻¹), and (vi) $Na_2S_2O_8$ (80 mmol L⁻¹) + $FeSO_4 \cdot 7H_2O$ (30 mmol L⁻¹). Five treatments were initiated for set (2): (i) control, (ii) $FeSO_4 \cdot 7H_2O$ (10 mmol L⁻¹) + $Na_2S_2O_8$ (60 mmol L⁻¹), (iii) $FeSO_4 \cdot 7H_2O$ (20 mmol L⁻¹) + $Na_2S_2O_8$ (60 mmol L⁻¹), (iv) $FeSO_4 \cdot 7H_2O$ (40 mmol L⁻¹) + $Na_2S_2O_8$ (60 mmol L⁻¹), and (v) $FeSO_4 \cdot 7H_2O$ (50 mmol L⁻¹) + $Na_2S_2O_8$ (60 mmol L⁻¹). For each set of experiment, 1 mL reagent solution was added in the following order: (1) sodium persulfate and (2) ferrous ion (one-fifth every 5 min) (Liang et al., 2004a). Afterwards, the sediments and solutions were vortex-mixed and incubated at ambient temperature (20 °C). Sampling and detection of the remaining amounts of DDTs were conducted at 0, 1, 2, 3, and 4 h. The concentration of sulfate in the supernatant was detected by ion chromatography at 4 h.

2.4 Extraction and Analysis of DDTs

The DDTs in the sediments were extracted using the ultrasonic extraction method. Hexane (10 mL) was added to the serum bottles containing the sediments. The bottles were then sonicated for 30 min in 6 L water in an ultrasonic bath (Soniclean, Australia) to ensure particle and solvent mixing (Thangavadivel et al., 2011). Furthermore, the contents were vortex-mixed and the solvent with DDTs was separated and passed through a funnel filled with 2.0 g anhydrous sodium sulfate (Na_2SO_4) to eliminate the remaining water in the samples (Hussen et al., 2006). Then, the extraction solvent with DDTs was evaporated to 2.0 mL using nitrogen gas.

The DDTs were analyzed by an Agilent 7890A gas chromatograph (Agilent Technologies, CA, USA) equipped with an Agilent 5975C mass selective detector (Agilent Technologies, CA, USA). The column used was HP-5MS, 30.00 m × 0.25 mm id with a film thickness of 0.25 μm. The injector temperature was 250 °C and the

helium gas flow rate was 1.0 mL min⁻¹. The column temperature was initially set to 40 °C for 1 min, and then later increased at a rate of 30 °C min⁻¹ to 130 °C min⁻¹. Afterwards, column temperature was switched to a rate of 5 °C min⁻¹ to 160 °C min⁻¹ for 6 min, and then, to a rate of 10 °C min⁻¹ to 190 °C for 3 min. Finally, the temperature was switched to a rate of 20 °C min⁻¹ to 300 °C and maintained isothermally for 2 min (Manirakiza et al., 2000). The injection volume was 1.0 µL in a splitless mode.

2.5 Analysis of the Sulfate in the Supernatant

According to Equations (1) and (2), persulfate will decompose into sulfate during reactions activated by ferrous ion. To evaluate the amount of decomposed persulfate, the concentration of supernatant sulfate was determined by ion chromatography. After 4 h reaction, 0.5 mL supernatant was passed through a 0.45 µm filter to eliminate impurities which may interfere with ion chromatography. The ion chromatography system, Dionex Ionpac AS14 column (4.6 mm × 3100.0 mm; Thermo Scientific, CA, USA), comprises a GP50 gradient pump, a column oven LC25, and an electrochemical detector ED50. Elution buffer was made of 3.5 mM sodium bicarbonate and 1.0 mM sodium carbonate. The flow ratio was 1.2 mL min⁻¹.

3. Results and Discussion

3.1 Assessment of DDT Contamination in Sampling Sites

Although the sampling site was a dicofol manufacturing factory which closed in 2002, no dicofol was detected in the samples because this pesticide is highly degradable in natural environments. However, DDT contamination remains serious even after nine years of production cessation. Dicofol impurity was proposed to possibly contribute to DDTs in the environment. This hypothesis was supported by the investigation of air samples collected over Taihu Lake, China during the summer of 2002, where very high concentrations of DDTs were found to be related to dicofol applications (Qiu et al., 2004), thus suggesting that dicofol is a possible source of DDTs which may constantly evaporate from soil to air.

According to the guidelines of the Chinese Environmental Quality Standard for Soil (GB15618-1995), the quality of soil can be classified as: with background pollution (grade I), low pollution (grade II), and high pollution (grade III). All three sampling sites were highly contaminated by DDTs, as shown in Table 1. *p,p'*-DDE and *p,p'*-DDT were the main DDT components and their concentration (1 mg kg⁻¹) was much higher than that of grade III soil. Moreover, *p,p'*-DDE concentration was higher than *p,p'*-DDT concentration in all samples. The results confirmed previous findings that DDE is the most abundant DDT component in soil and sediments (Guo et al., 2009), and is hardly degraded compared with other DDT components (Thomas et al., 2008; de la Cal et al., 2008). Further investigations and an effective remediation option for dicofol-type DDT contamination are recommended.

Table 1. Concentrations of DDTs in samples obtained from the three sampling sites

Station	Orientation	<i>p,p'</i> -DDT (mg kg ⁻¹) (Mean±SE)	<i>p,p'</i> -DDE (mg kg ⁻¹) (Mean±SE)
No. 1	N: 39°14'26.1" E: 117°06'33.7"	4.38±0.45	22.99±1.21
No. 2	N: 39°14'23", E: 117°06'34.14"	10.27±1.54	14.34±1.61
No. 3	N:39°14'23.3", E:117°06'33.7"	11.84±0.97	115.27±7.52

The samples were collected in 2010. Mean ± SE values (mg kg⁻¹) are shown (n = 3). Means are significantly different (one-way ANOVA: p < 0.05).

3.2 Influence of S₂O₈²⁻ Concentration on DDT Degradation

To investigate the effect of persulfate contents on DDT degradation at Fe²⁺ concentration of 30 mM at ambient temperature (20 °C), different concentrations of persulfate were used. For all S₂O₈²⁻/Fe²⁺ molar ratios, DDT degradation and persulfate decomposition (represented by sulfate formation) were observed (Figures 1 and 2). DDT degradation occurred almost instantaneously and stabilized because of the high reactive activity and unstability of SO₄⁻ generated from S₂O₈²⁻-Fe²⁺ reaction. For *p,p'*-DDE, the remaining ratios after 4 h decreased

from 71% to 40% at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 10/30 to 80/30, respectively (Table 2). An increase in persulfate concentrations when Fe^{2+} level is 30 mM resulted in increased DDT degradation. However, an insignificant increase in DDT degradation was observed when $S_2O_8^{2-}/Fe^{2+}$ molar ratio was higher than 60/30. The *p,p'*-DDT remaining ratios after 4 h were 23%, 19%, 25%, 4%, and 5% at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 10/30, 20/30, 40/30, 60/30, and 80/30, respectively (Table 2). The influence of various persulfate concentrations on the degradation time course of DDTs (*p,p'*-DDE and *p,p'*-DDT) (Figure 1) indicated that total degradation ratios (in ascending order) were $10/30 < 20/30 < 40/30 < 60/30 < 80/30$. Increasing the amount of persulfate content did result in an increase in sulfate formation and a proportional increase in DDT degradation except for the 80/30 $S_2O_8^{2-}/Fe^{2+}$ treatment (Figure 2). The concentration amounts of sulfate originating from the decomposition of persulfate were 10.04, 10.14, 10.28, 12.64, and 12.83 $mg\ g^{-1}$ at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 10/30, 20/30, 40/30, 60/30, and 80/30, respectively. The amount of DDT degradation were 43, 52, 71, 89, and 91 $\mu g\ g^{-1}$ at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 10/30, 20/30, 40/30, 60/30, and 80/30, respectively. The $S_2O_8^{2-}/Fe^{2+}$ molar ratio of 60/30 is the most effective and economical proportion. The observed optimal reaction ratio of $S_2O_8^{2-}/Fe^{2+}$ is greater than the theoretical stoichiometric ratio of 1:1 according to Equation (1). Persulfate anions were depleted by several reactions, such as organic substance oxidation, apart from those activated by Fe^{2+} (Kolthoff et al., 1951). The result was similar to that reported by Liang et al. (2004a) who suggested that the $S_2O_8^{2-}/Fe^{2+}$ molar ratio of 40/30 is the most effective proportion for TCE degradation in aqueous systems, and that beyond this ratio, adding persulfate only resulted in an insignificant increase in TCE degradation. In this study, an excess concentration of persulfate was necessary because persulfate was consumed by sediment materials. As for $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 10/30, 20/30, and 40/30, the initial persulfate concentrations were insufficient when considering the stoichiometric ratio of reactions between persulfate and ferrous ions. Without adding persulfate and ferrous ion, no DDT degradation was observed for 4 h.

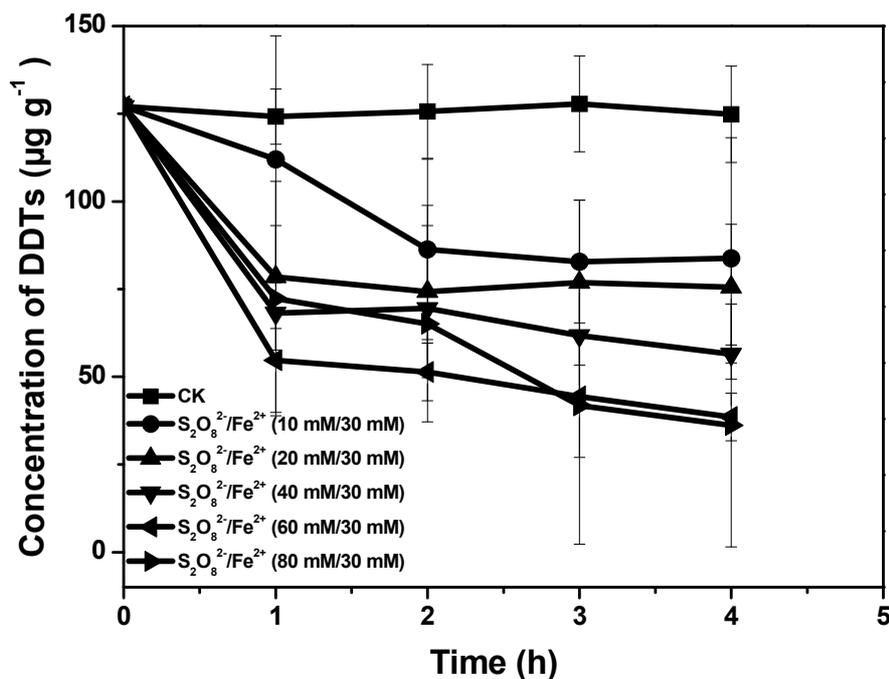


Figure 1. Dynamics of DDT degradation with different persulfate concentrations

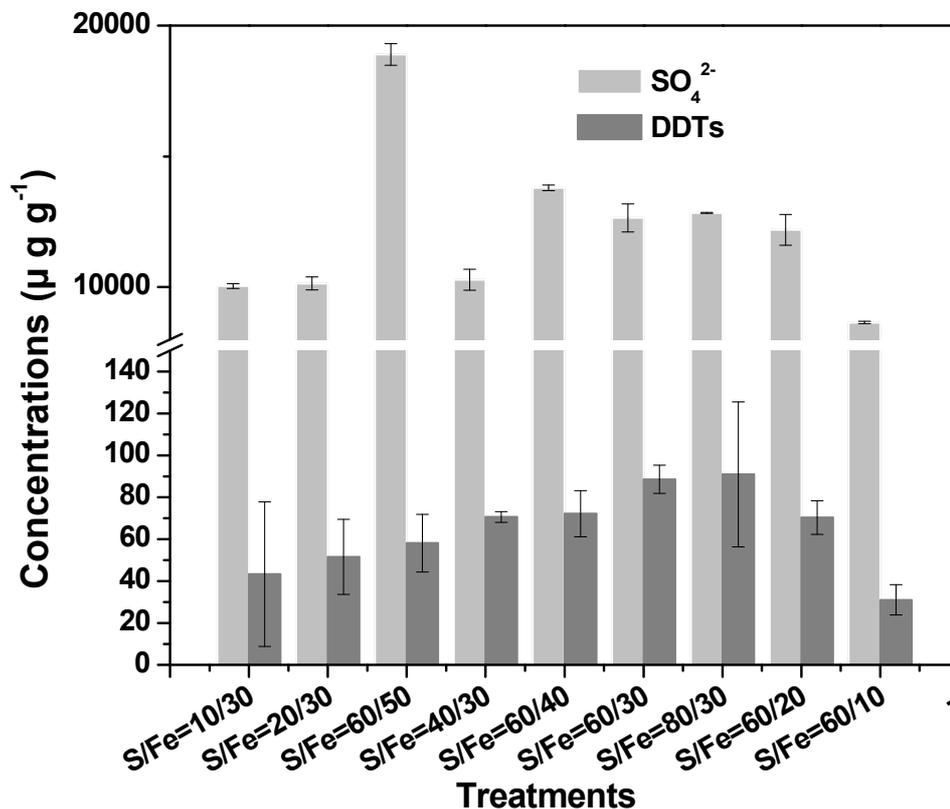


Figure 2. DDT degradation and sulfate formed by persulfate decomposition after 4 h

Table 2. The influence of different persulfate concentrations on the degradation of (A) DDE and (B) DDT

S ₂ O ₈ ²⁻ /Fe ²⁺ molar ratio	Degradation ratio (C _t /C ₀ %)	
	DDE (Mean±SD)	DDT (Mean±SD)
10/30	29.06±11.02	76.95±29.30
20/30	36.21±9.20	80.84±23.00
40/30	50.67±2.55	74.93±3.17
60/30	64.09±15.64	95.77±20.35
80/30	39.94±35.14	94.66±86.32

3.3 The Influence of Fe²⁺ Concentration on DDT Degradation

To further elucidate the effect of available ferrous ion at ambient temperature (20 °C) on the degradation of DDTs, ferrous ion was added in one-fifth increments to the reaction bottles containing S₂O₈²⁻ concentration of 60 mM at 5 min intervals. Figure 3 shows the effect of different Fe²⁺ concentrations on DDT degradation and persulfate decomposition. DDT degradation and persulfate decomposition (indicated by sulfate formation) were observed in all S₂O₈²⁻/Fe²⁺ molar ratios. After five successive additions of Fe²⁺, the initial S₂O₈²⁻/Fe²⁺ molar ratios were 60/10, 60/20, 60/30, 60/40, and 60/50. The remaining *p, p'*-DDE after 4 h were 100%, 63%, 36%, 60%, and 59%, respectively (Table 3). In the case of the 60/10 S₂O₈²⁻/Fe²⁺ molar ratio, *p, p'*-DDE degradation was insignificant because the concentration of Fe²⁺ was insufficient. Increasing Fe²⁺ concentration from a S₂O₈²⁻/Fe²⁺ molar ratio of 60/10 to 60/30 resulted in approximately 64% increase in *p, p'*-DDE degradation. Theoretically, increasing Fe²⁺ concentration would produce more SO₄⁻ which could promote DDT degradation. However, *p, p'*-DDE degradation decreased when S₂O₈²⁻/Fe²⁺ molar ratio was higher than 60/30. Similar phenomena have been reported by other researchers (Kislenko et al., 1995; Liang et al., 2004a). The possible consumption of sulfate free radicals may result from reactions with H₂O, S₂O₈²⁻, and excess Fe²⁺ (Kolthoff et al.,

1951; Mcelroy & Waygood, 1990), as shown in Equations (2), (3), and (4).



The remaining *p, p'*-DDT after 4 h were 4.92%, 3.29%, 4.23%, 6.07%, and 2.29% at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 60/10, 60/20, 60/30, 60/40, and 60/50, respectively (Table 3). The rates of *p, p'*-DDT degradation were all significant whether Fe^{2+} content was sufficient or insufficient. Fe^{2+} has also been reported to result in significant *p, p'*-DDT reductive degradation (Li et al., 2010). The observed $S_2O_8^{2-}-Fe^{2+}$ system in this study showed that the competition for $SO_4^{\cdot-}$ between excess Fe^{2+} and DDTs was significant (Figure 2). Further increases in Fe^{2+} contents resulted in proportional increases in persulfate decomposition but not in DDT degradation (Figure 2). The sulfate obtained from the decomposition of persulfate, accompanied by increasing Fe^{2+} amounts, were 8.64, 12.19, 12.64, 13.79, and 18.89 $mg\ g^{-1}$ at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 60/10, 60/20, 60/30, 60/40, and 60/50, respectively. The rates of DDT degradation were 31, 70, 89, 72, and 69 $\mu g\ g^{-1}$ at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 60/10, 60/20, 60/30, 60/40, and 60/50, respectively. In this case, the rates of DDT degradation were not in agreement with the sulfate concentrations, thus indicating that high amounts of Fe^{2+} resulted in a competition for $SO_4^{\cdot-}$ between excess Fe^{2+} and DDTs at $S_2O_8^{2-}/Fe^{2+}$ molar ratios less than 60/30. Moreover, the degradation time course of DDTs (*p, p'*-DDE and *p, p'*-DDT) (Figure 3) indicated that total degradation ratios (in ascending order) were 60/10 < 60/50 < 60/20 < 60/40 < 60/30. Therefore, sulfate free radicals formed through ferrous-ion activation are capable of effective DDT oxidative degradation, and that the optimum molar ratio for $S_2O_8^{2-}/Fe^{2+}$ was 60/30.

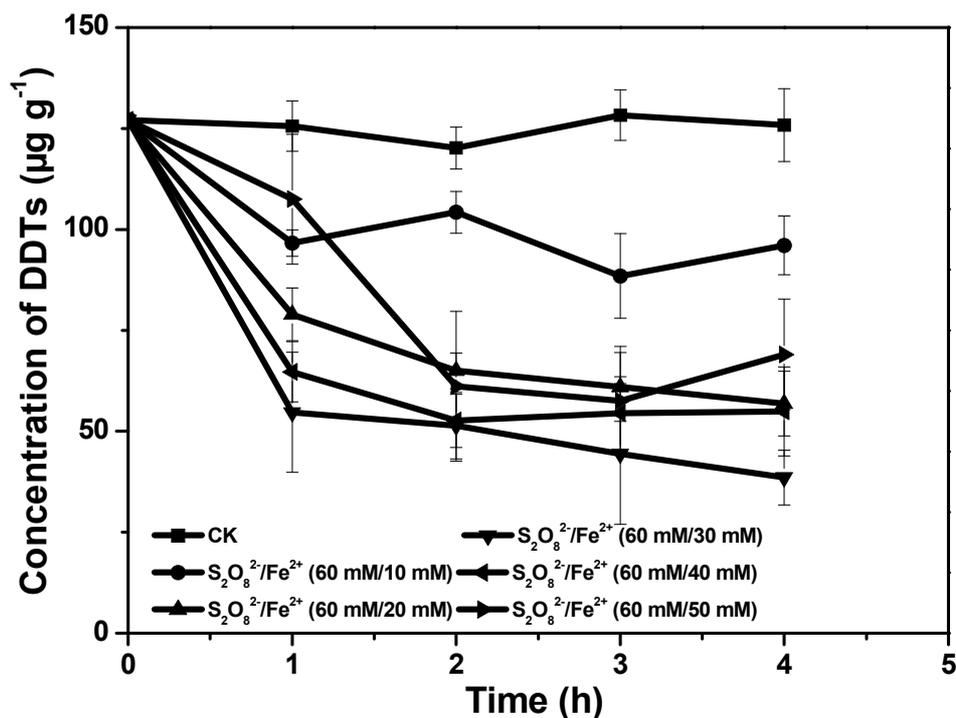


Figure 3. Time course of DDT degradation with different concentrations of ferrous ion

Table 3. Influence of different ferrous ion concentrations on the degradation of (A) DDE and (B) DDT

$S_2O_8^{2-}/Fe^{2+}$ molar ratio	Degradation ratio (C_t/C_0 %)	
	DDE (Mean±SD)	DDT (Mean±SD)
60/10	0	95.08±7.56
60/20	37.37±6.21	96.72±13.20
60/30	64.09±13.90	95.77±21.21
60/40	39.98±8.07	93.93±16.98
60/50	40.14±9.96	97.71±13.32

4. Conclusions

The assessment of the site for a dicofol manufacturing factory showed serious DDT (*p*, *p'*-DDE and *p*, *p'*-DDT) contamination in sediments. DDT degradation and persulfate decomposition were observed by calculating the amount of ferrous ion and persulfate at ambient temperature (20 °C). This study demonstrated that sulfate free radicals formed by ferrous ion activation are capable of degrading DDTs in sediments. Total DDT degradation ratios at $S_2O_8^{2-}/Fe^{2+}$ molar ratios of 31, 43, 52, 69, 70, 71, 72, 89, and 91 $\mu\text{g g}^{-1}$ were 60/10 < 10/30 < 20/30 < 60/50 < 60/20 < 40/30 < 60/40 < 60/30 < 80/30, respectively. However, ferrous ion would react with SO_4^- and DDT degradation efficiency would decrease if ferrous ion contents became excessive. Further increases in persulfate concentration beyond an $S_2O_8^{2-}/Fe^{2+}$ molar ratio of 60/30 resulted in insignificant increases in DDT degradation. Thus, an $S_2O_8^{2-}/Fe^{2+}$ molar ratio of 60/30 was the most effective and economical proportion. This result indicated that a slow and steady production of free radicals is most desirable, and that Fe^{2+} availability plays an important role in controlling persulfate reactions activated by ferrous ion. Fe^{2+} -activated persulfate may be significant in developing an effective and environment friendly remediation option for DDT-contaminated sediments and soils. However, further work is necessary to determine how practical applications can be accomplished.

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